## What is the change in velocity in this elastic collision?



What is the change in momentum of particle in this elastic collision?



## How many particles moving with velocity **v** will hit a wall in time $\Delta t$ ?



Only those close enough to the wall.



Only those close enough to the wall. And close enough is  $d = v\Delta t$ . How many particles will hit a wall in time  $\Delta t$ ?



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Only those close enough to the wall. And close enough is  $d = v\Delta t$ . How many particles are in the entire box? Let's call that number *N*. That's *N* particles in volume *AL*.



How many are close enough to the wall to hit it in time  $\Delta t$ ? Let's call that number *x*. That's *x* particles in volume  $A(v\Delta t)$ . N particles are in the entire volume.

x particles are in the volume which is close enough for particles to hit the wall in time  $\Delta t$ .

The number of particles per unit volume is the same for both the entire volume and the smaller volume.



x particles are in the volume that is close enough for particles to hit the wall in time  $\Delta t$ .

 $x = \frac{N}{V}A(v\Delta t)$ 



But how many particles in that little volume will actually hit the wall. They are close enough, but not all of them are actually moving *toward* the wall. We will now rely on statistics. There are 6 directions for a molecule to move in. (Can you imagine them?) One of those directions is toward the wall. So, on the average, 1/6 of the x particles will actually hit the wall in time  $\Delta t$ . x particles are in the volume which is close enough for particles to hit the wall in time  $\Delta t$ . Only 1/6 of those are heading toward the wall.

We will use the symbol *n* for the number of particles that actually hit the wall in time  $\Delta t$ :





 $P = \frac{1}{3} \frac{N}{V} (mv^2)$  $P = \frac{2}{3} \frac{N}{V} \left(\frac{1}{2} m v^2\right)$  $PV = \frac{2}{3}N\left(\frac{1}{2}mv^2\right)$ PV = NkTPV = nRT $NkT = \frac{2}{3}N\left(\frac{1}{2}mv^2\right)$  $\frac{3}{2}kT = \frac{1}{2}mv^2$